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~~Reduction Reactions Nature of Oxidation and Reduction Questions 1.~~

~~Explain in your own words the following oxidation processes: a. Gain~~

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~~own words the following reduction processes: a. Loss of oxygen b.~~

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~~Solids In your textbook, read about liquids and solids. In the space at~~

~~the left, write true if the statement is true; if the statement is false,~~

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704 CHAPTER 22 Identify the product that balances the following
nuclear reaction: ${}_{84}^{212}\text{Po} \rightarrow {}_{2}^{4}\text{He} + \text{?}$. The total mass number and

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atomic number must be equal on both sides of the equation. $212\ 84\text{Po}$
 $42\text{He} +$ mass number: $212 - 4 = 208$ atomic number: $84 - 2 = 82$ 2.
The nuclide has a mass number of 208 and an atomic number of
 $82, 208\ 82\text{Pb}$. 3.

CHAPTER 22 Nuclear Chemistry

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NUCLEAR CHEMISTRY 705 SECTION 22-2 OBJECTIVES Define
and relate the terms radioactive decay and nuclear radiation Describe
the different types of radioactive decay and their effects on the nucleus
Define the term half-life, and explain how it relates to the stability of a
nucleus Define and relate the terms decay series, parent nuclide ...

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Use dimensional analysis and the given densities to make the following conversions. a. 14.8 g of boron (B) to cubic centimeters of boron. The density of boron is 2.34 g/ cm³. $(14.8 \text{ g}) \left(\frac{1 \text{ cm}^3}{2.34 \text{ g}} \right) = 6.32 \text{ cm}^3$ b. 2.8 L of argon (Ar) to grams of argon.

~~Chapter 4 Problem Solving in Chemistry Answer Key~~

1) No, it is not a proper answer; you do not know whether the professor meant homework problem number 20 or 20 homework

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problems. Twenty is a quantity and does not make sense without a unit. 3) Middle Zeros: Zeros that appear between other nonzero digits are ALWAYS significant.

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