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[Section 8.4 – Polar Bonds and Molecules.](#) [Covalent bonds involve sharing electrons between atoms. When the atoms in the bond pull equally, the bonding electrons are shared equally, and the bond is nonpolar. When the atoms in the bond pull unequally, the bonding electrons are pulled closer to one atom, and the bond is polar.](#)

Chapter 8 – Covalent Bonding

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[The sp 3 orbital overlaps with hydrogen's 1s orbital to form a covalent bond, called ? bond. The angle between each lobe assumes 109.47° since the sp 3 carbon is tetrahedral. sp 2 Hybrid Orbitals : If a carbon atom is in a double-bond environment, such as C 2 H 4 \(ethylene\), the carbon is hybridized to become sp 2 hybrid orbital.](#)

Chapter 8. Covalent Compounds: Bonding Theories and ...

[Lattice energies of alkali metal iodides.](#) [Copyright McGraw-Hill 2009 11.](#) [The ionic radii sums for LiF and MgO are 2.01 and 2.06 Å, respectively, yet their lattice energies are 1030 and 3795 kJ/mol.](#)

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[In covalent bonds, electron sharing usually occurs so that atoms attain the electron configurations of noble gases. • For example, a single hydrogen atom has one electron. But a pair of hydrogen atoms shares electrons to form a covalent bond in a diatomic hydrogen molecule.](#)

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