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Naming Compounds  
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Writing Formulas with  
Polyatomic Ions

Binary Ionic  
Compounds Ionic  
Formulas Answers

Write the ionic equation  
for the word equation.

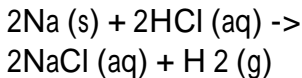
Sodium (s) +  
hydrochloric acid (aq)  
-> sodium chloride (aq)



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+ hydrogen (g) Solution:

Step 1: Write the equation and balance it.



Step 2: Split the ions.

(Only compounds that are aqueous are split into ions.)

Writing Ionic Equation  
(video lessons, examples  
and solutions)

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Write the chemical formula for an ionic compound composed of each pair of ions. the calcium ion and the oxygen ion; the  $2+$  copper ion and the sulfur ion; the  $1+$  copper ion and the sulfur ion;

Answer a:  $\text{CaO}$  Answer

b:  $\text{CuS}$  Answer c:  $\text{Cu}_2\text{S}$

5.5: Writing Formulas for Ionic Compounds -

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Worksheet ...

Answer key ionic  
compound formula  
writing worksheet write  
chemical formulas for the  
compounds in each box  
the names are found by  
naming ionic  
compounds answer key

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Formulas  
Answers

give the name and molar mass of the following ionic compounds name 1  
na 2 co 3 sodium carbonate 2 naoh sodium.

Writing Formulas Ionic Compounds Chem  
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want to supply all you need a single real and trusted origin, most of us current helpful info on several subject matter in addition to topics. By suggestions about conversation crafting, to developing publication sets out, or even to pinpointing which type of ...

Writing Ionic Formulas

*Page 15/31*

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Worksheet Answers |  
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2 Li<sup>+</sup> ions are required  
to balance 1 O<sup>2-</sup> ion.

Nickel has a charge of +2  
and sulfur has a -2  
charge, therefore. 1 Ni<sup>2+</sup>  
ion is required to balance  
1 S<sup>2-</sup> ion. Bismuth has a  
+3 charge and Fluorine  
has a -1 charge, therefore.  
1 Bi<sup>3+</sup> ion is required to  
balance 3 F<sup>-</sup> ions.



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Predicting Formulas of  
Ionic Compounds  
Example Problem

Binary Ionic Formula

Practice Name\_\_\_\_\_ 1)

sodium chloride  $\text{Na}^{+1}$

$\text{Cl}^{-1}$   $\text{NaCl}$  2) lithium

bromide  $\text{Li}^{+1}$   $\text{Br}^{-1}$   $\text{LiBr}$

3) magnesium fluoride

$\text{Mg}^{+2}$   $\text{F}^{-1}$   $\text{MgF}_2$  4)

potassium oxide  $\text{K}^{+1}$

$\text{O}^{-2}$   $\text{K}_2\text{O}$  5) calcium

sulfide  $\text{Ca}^{+2}$   $\text{S}^{-2}$   $\text{CaS}$  6)

aluminum iodide  $\text{Al}^{+3}$

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I-1 All3 7) barium  
bromide  $\text{Ba}^{+2} \text{Br}^{-1}$   
 $\text{BaBr}_2$

Binary Ionic Formula

Practice - LSRHS

From the following ionic  
compounds, choose the  
name-formula pair that is  
not correctly matched.

answer choices. sodium  
sulfide  $\text{Na}_2\text{S}$ . ammonium  
nitrate.  $\text{NH}_4\text{NO}_3$ .  
sodium sulfate.  $\text{Na}_2\text{SO}_3$ .

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Formulas  
Answers  
calcium oxide  $\text{CaO}$ .

$\text{Na}_2\text{S}$   
>.

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## Ionic Formulas.

Chemical formulas for ionic compounds are called ionic formulas. A proper ionic formula has a cation and an anion in it; an ionic compound is never formed between two cations only or two anions only. The key to writing proper ionic formulas is simple: the total positive charge must balance the total negative

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## Answers

3.4: Ions and Ionic  
Compounds - Chemistry  
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3 -2OH-SO. 4 -2PO. 4  
-3NO. 3 -. Na+NaCl.  
Na. 2CO. 3NaOH Na.  
2SO.

Writing Formulas (KEY)  
(Criss-Cross Method)  
Ionic compounds are  
compounds made up of

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ions. An ionic compound is formed by the complete transfer of electrons from a metal to a nonmetal, and the resulting ions have achieved an octet. Below is a chemistry quiz on ionic compounds, names, and formulas, give it a shot and see if you understood all we covered in this topic on Ions.

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Chemistry Quiz: Ionic  
Compound Names And  
Formulas ...

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compounds when given  
the formula. Practice  
naming ionic  
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Naming ionic compounds is an important skill in chemistry. This is a collection of ten chemistry test questions dealing with naming ionic compounds and predicting the chemical formula from the

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compound name. The answers are at the end of the test.

## Naming Ionic Compounds Test Questions - ThoughtCo

I. Name the following Ionic Compounds:!

1. LiBr lithium bromide !
2. !  $\text{CuC}_2\text{H}_3\text{O}_2$  ! copper (I) acetate !
3. !  $\text{PbSO}_3$  ! lead (II) sulfite !
4. !  $\text{NaClO}_3$  ! sodium

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chlorate ! 5. !  $\text{CaC}_2\text{O}_4$  !  
calcium oxalate ! 6. !  
 $\text{NaHSO}_4$  ! sodium  
bisulfate ! 7. !  $\text{Hg}_2\text{Cl}_2$  !  
mercury (I) chloride ! 8. !  
 $\text{NaHCO}_3$  ! sodium  
bicarbonate ! 9. !  $\text{NiBr}_3$  !  
nickel (III) bromide ! 10.  
!  $\text{AuCl}_3$  ! gold (III)  
chloride ! 11. !  $\text{KMnO}_4$  !

Ionic Bonding and Writing Formulas Name: KEY  
!!! Part A. Use ...

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In writing the formula of an ionic compound, the net ionic charge must be

\_\_\_\_\_. Zero. The

systematic name for

baking soda ( $\text{NaHCO}_3$ )

is a sodium bicarbonate.

NT. In writing a formula

for an ionic compound,

the net ionic charge of

the formula must be zero.

AT. Anions that contain

oxygen end in -ite or

-ate. ST.

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