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occurs when two water molecules produce
a hydronium (H_3O^+) and a hydroxide
(OH^-) ion $2 \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ The Page
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ionization of water occurs when two water
molecules produce a hydronium (H_3O^+)
and a hydroxide (OH^-) ion $2 \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ +$

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OH- The ionization constant (K_w) of water is: $K_w = [H_3O^+] [OH^-] = 1.0 \times 10^{-14} M$ Acidic, Basic, & Neutral IF $[H_3O^+] >$

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CHAPTER 15 REVIEW Acid-Base

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Titration and pH SECTION 2 SHORT

ANSWER Answer the following questions in the space provided. 1. Below is a pH curve from an acid-base titration. On it are labeled three points: X, Y, and Z. a. Which point represents the equivalence point? b. At which point is there excess acid in the system? c.

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~~15 Acid-Base Titration and pH~~

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mhardie11. Terms in this set (10) Self-

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produce a hydroniumion and a hydroxide

ion by transfer of a proton. pH. the

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negative of the common logarithm of the hydronium ion concentration. pOH .

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Titration and pH CHAPTER 15 REVIEW

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~~Review Answers~~ Acid-Base Titration and pH SECTION 2
SHORT ANSWER Answer the following questions in the space provided. 1. Below is a pH curve from an acid-base titration. On it are labeled three points: X, Y, and Z.
_____ a.

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