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2 POGIL™ Activities for High School Chemistry 3. In the space under each reaction in Model 1, write the oxidation number for every atom. Divide the work among your group members. An example is shown here:  $4\text{Fe}(s) + 3\text{O}_2(g) \rightarrow 2\text{Fe}_2\text{O}_3(s)$  0 0 +3 -2 0 +3 -2

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2.4. Identify any elements that changed oxidation number in the reactions in Model 1.

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1. Consider the incomplete half-  
reactions below. a. Use oxidation  
numbers to identify the reactions  
below as oxidation or reduction. b.  
Place the correct number of electrons  
on the appropriate side of the reaction  
to complete the equation.  
PEOVCTI&..J 6 x STOP CF. e- 12.  
Consider Reaction A in Model 2.

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