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Molarity Answers

Many companies, including pharmaceutical companies, keep track of the concentration of a solution by measuring its molarity — a ratio of the number of solute particles to the volume of the solution. In this activity you will learn about molarity and how to represent

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quantitatively. Model 1
□ Lemonade Mixtures*

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chemists is molarity (M). Molarity is defined as the number of moles of solute (mol) divided by the total volume (V) of the solution in liters (L). Molarity = moles of solute per liter of solution ($M = \text{mol} / \text{L}$).

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of the new solution?
(Answer: 9-801 HN03)

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Answer: 0.0200 M

3. 30.00 ml, of water is added to 20.00 ml, of 0.0500 M sodium carbonate. What is the molarity of sodium carbonate in the new solution? (Answer: 0.0200 M)

4. What volume of 0.100 M NaCl is needed to make 5.00 ml, of 0.200 M solution? (Answer: 0.10 ml)

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Solutions and Molarity

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Answer Key Chemistry

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